



Chemical Modification of Polystyrene Merrifield: Extraction of Zinc and Magnesium Located in Wastewater

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Abstract: In order to remove metals (Zn & Mg) located in wastewater, a new series of clickable polystyrene Merrifield grafted with azide and alkynes were synthesized and implicated in this domain. In a first step, we transformed the polystyrene Merrifield into the known Azidomethyl polystyrene, in a second step, a coupling reaction between the Azidomethyl polystyrene and dipropargylamine was realized for the synthesis of a new terminated polystyrene alkyne containing one unit of (1,4)-triazole. To increase the length of the chain grafted on the polystyrene Merrifield, several coupling reactions are applied, the main idea of this work was to increase the number of (1,2,3) –triazole units of this grafted chain. The click chemistry based on the Huygens's reaction catalyzed by copper (I) was used in this synthesis, the new polymers containing (1,4)-disubstituted triazole are tested for the extraction of Mg and Zn located in wastewater. The structure of new obtained polymers was confirmed by infrared spectroscopy (ATR-FTIR), UV-visible spectroscopy and elemental analysis for nitrogen. Differential Thermal Analysis (DTA) and Thermo Gravimetric Analysis (TGA) were used to study crosslinking behavior of these polymers. This study shows a low selectivity of studied polymers for the retention of magnesium, while for the Zinc, the percentage removal was average, it's in the order of 30%.

Keywords: Polystyrene Merrifield, (1,4)-Triazoles, Extraction, Zinc and Magnesium in Wastewater

1. Introduction

Water was an important source of sustainable development, it was used for the fight against poverty, economic growth and environmental sustainability. For this reason, the purification of wastewater was the recent idea in this work. In our day, pollution has become an important problem, because a major part of toxic pollutants was introduced into urban wastewater. Heavy metals (MX_n) [1, 2], organic compounds (O. C) [3] and nitrate ions (NO₃⁻) [4] constituted the major toxic pollutants rejected in urban wastewater. The elimination of these pollutants from the aquatic environment was not usually achieved by natural processes, because these pollutants established between them a stable complexes. Farmers used

wastewater in irrigation, because they were rich in nutrients elements [5]. The pollutants located in wastewater can be accumulated in the soil, can be absorbed by plants [6, 7], and can be also percolated from groundwater [8]. The toxicity of heavy metals is mainly due to their tendency to accumulate in living organisms and to concentrate along trophic chains [9]. The objective of this work was the synthesis of new polymers containing (1,4)-triazole units from polystyrene Merrifield and testing them in the elimination of (Zn and Mg) localized in wastewater. The click chemistry based on the Huygens's reaction catalyzed by copper (I) iodide was used in this synthesis [10-14]. Figure 1 summarized the approach of this work.

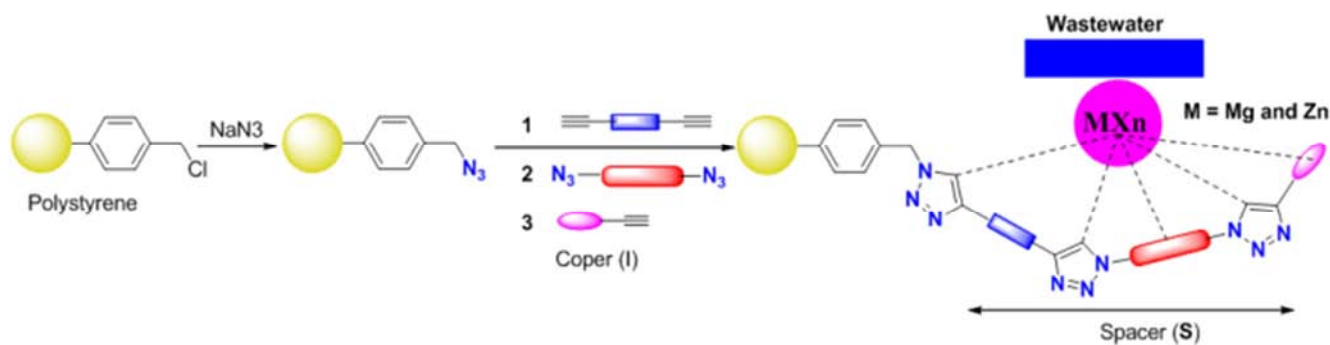


Figure 1. General scheme for synthetic (1,4)-triazoles from Polystyrene Merrifield and its application in extraction of (Zn & Mg) located in wastewater.

2. Experimental

2.1. Materials

Merrifield polymer (chloromethylated poly-styrene-co-divinyl-benzene (1%), 200–400 mesh, 1.1 mmol Cl.g⁻¹) was used as-received from Fluka, 1-chloro-2-[2-(2-chloroethoxy) ethoxy] ethane (C₆H₁₂O₂Cl₂, Mw = 187.065; 99%), Quinic acid (C₇H₁₂O₆, Mw = 192.17g.mol⁻¹), propargylamine (C₃H₅N, Mw = 55.08g.mol⁻¹; 98%), di-propargyl amine (C₆H₇N, Mw = 93.13g.mol⁻¹; 97%), copper (I) iodide (99,5%), sodium azide (99%), sodium sulfate (99%), acetone (97%), N,N-dimethyl formamide (DMF), diethyl ether (99%), methylene chloride (CH₂Cl₂), Tetrahydrofuranne (THF), Dimethyl sulfoxide (DMSO), Methanol (MeOH) and tri-ethylamine (ET₃N) were purchased by Sigma Aldrich. The samples of wastewater were taken from the station of treatment of Beja (36° 44' North, 9° 11' East, Tunisia).

The collection of wastewater was taken in the first week of January 2017. The sampling of wastewater was taken by immersing the bottle with 1L of capacity in the basin at 15cm in depth. The sample was maintained at 4°C and filtered before treatment by polymers.

Infrared analysis using the attenuated total reflectance technique (ATR/FTIR) was made on a Nicolet IR 200 FTIR spectrometer between 4000 and 400cm⁻¹. The relevant absorption bands of vibrations (ν) was indicated. Elemental analysis of (N) was performed by using Perkin Elmer Analyzer CHN Series II 2400. Nuclear magnetic resonance

spectroscopy (NMR) was recorded on a Bruker Avance 300 at 300 MHz for the proton spectra. Differential thermal analysis DTA was performed on a SETARAM TGA92, Thermal Gravimetric Analysis was performed on a Labsys TG. The samples (10-15mg) were placed in 40 μ L aluminum pans and analyzed under argon between 25 and 600°C at a heating speed of 10°C.min⁻¹. The UV-Vis spectra were carried out using UV-Vis spectrophotometer BKUV 1600 PC in the DMF, atomic absorption spectroscopy (AAS) analysis (Perkin-Elmer A Analyst 200) operating with an air-acetylene flame was used to measure residual metal ion concentrations.

2.2. Organic Synthesis

2.2.1. Synthesis of

1-azido-2-[2-(2-azidoethoxy)ethoxy]ethane (A)

Sodium azide (5.85g, 90mmol) was added to a solution of 1-chloro-2-[2-(2-chloroethoxy)ethoxy]ethane (5.61g, 30mmol, 33%) in 50mL of DMF. The reaction mixture was stirred under nitrogen at 80°C for 20h. The resulting product was washed with distilled water (750mL) and extracted by diethyl ether (3 x 50mL). The organic phase was washed three times with distilled water (3 x 50mL) for removing unreacted sodium azide, and dried over MgSO₄. After filtration, the solvent was evaporated under vacuum. The resulting compound was colorless viscous (4.18g, 21mmol, 70%) (Figure 2).

¹H NMR, (300 MHz, CDCl₃, δ): 3.63 (t, 4H); 3.67 (t, 4H); 3.78 (t, 4H).

ATR/FTIR: 2873 (ν_{CH_2}), 2094 (ν_{N_3}), 1288 ($\nu_{\text{C-N}}$) cm⁻¹.

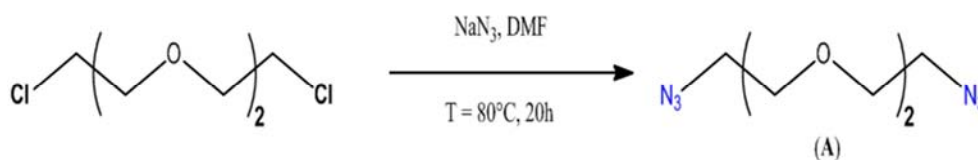


Figure 2. Reaction scheme of synthesis of 1-azido-2-[2-(2-azidoethoxy)ethoxy]ethane (A).

2.2.2. Synthesis of (1S, 3R, 5R, 7R)-5,7-dihydroxy-2,2-dimethyl-N-(prop-2-yn-1-yl)hexahydro-1,3-benzodioxole-5-carboxamide (C)

The alkyne (1S, 3R, 5R, 7R)-5,7-dihydroxy-2,2-dimethyl-N-(prop-2-yn-1-yl)

hexahydro-2H-1,3-benzodioxole-5-carboxamide was prepared using the method taken from the literature [15]. The steps of this synthesis were presented in the Figure 3.

¹H NMR, B, C₁₀H₁₅O₅: (300 MHz, CDCl₃, δ): 1.36 (s, 3H), 1.55 (s, 3H), 2.23 (d.d, J=15 Hz, J=3 Hz, 1H), 2.40 (m, 2H), 2.68 (dd, J=15 Hz, J=3 Hz, 1H), 3.71 (s, 1H), 4.31-4.35 (m,

1H), 4.51-4.54 (m, 1H) and 4.72-4.75 (m, 1H) ppm.

¹HNMR, C, C₁₃H₁₉NO₅: (300 MHz, CDCl₃, δ): 1.26 (s, 3H), 1.43 (s, 3H), 1.91-2.30 (m, 5H), 3.41 (s, 1H), 3.86-4.12 (m, 4H), 4.53-4.55 (m, 1H) and 7.41 (s, 1H) ppm.

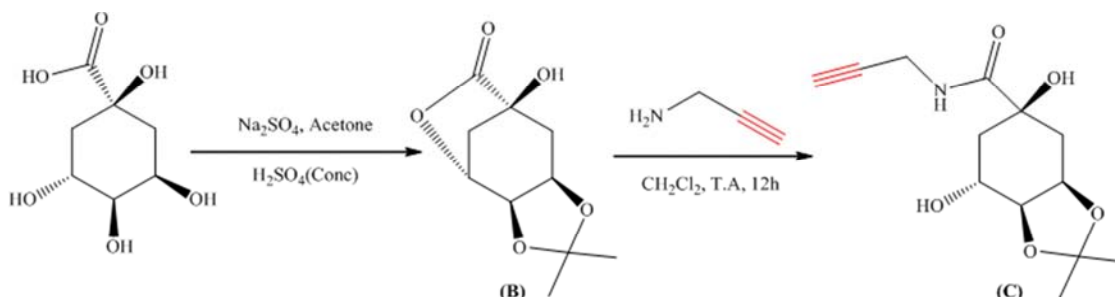


Figure 3. Reaction scheme of synthesis of (1S, 3R, 5R, 7R)-5,7-dihydroxy-2,2-diméthyl-N-(prop-2-yn-1-yl) hexahydro-1,3-benzodioxole-5-carboxamide (C).

ATR/FTIR, B: 2984 (ν_{C-H}), 1747 ($\nu_{C=O}$), 1051 (ν_{C-N}), 1122 (ν_{C-O}) cm^{-1} .

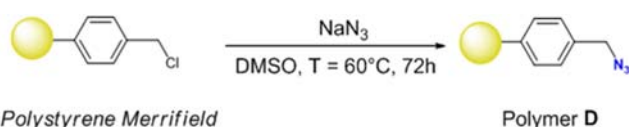
ATR/FTIR, C: 3417 (ν_{O-H}), 3286 (ν_{N-H}), 2985-2928 (ν_{CH}), 1653 ($\nu_{C=O}$), 1043 (ν_{C-N}) cm^{-1} .

2.2.3. Synthesis of Azidomethyl Polystyrene (D)

The method of preparation of Azidomethyl polystyrene (D) presented in Figure 4 was described by Ouerghui *et al.* [16]. The weight of Merrifield polymer (1.1 mmol Cl.g⁻¹) used in this synthesis was 10g, the weight of polymer azide obtained was 9.5g.

Element analysis: %(N)=3.726 or 2.66 N.g⁻¹ which give 0.89mmol N₃.g⁻¹, the yield of the substitution reaction was 81%.

ATR/FTIR: 3034 (ν_{Car-H}), 2929 (ν_{CH_2}), 2098 (ν_{N_3}), 1022 (ν_{C-N}) cm^{-1} .



Polystyrene Merrifield

Polymer D

Figure 4. Reaction scheme of synthesis of Azidomethyl polystyrene (D).

2.2.4. Synthesis of [4-(N-propynylamino Methyl)Triazol-1-yl]Methyl Polystyrene (E)

5g of Azidomethyl polystyrene (D: 4.45mmol of N₃) were suspended in 85mL of THF, after a low agitation for 30 minutes at room temperature, 0.589g of di-propargyl-amine (6,325mmol, 1.4eq), 0.026g of copper iodide (0,14mmol, 3%) and 8mL of ET₃N (5.81g, 57.5mmol, 13eq) were added. The suspension was slowly stirred at T=40°C for 72 hours, After filtration under vacuum, the obtained polymer was washed three times with THF (30mL), ET₃N (10mL), methylene chloride (10mL) and MeOH (10mL) alternatively and then dried at 40°C for 48 hours, the weight of new isolated polymer after purification was 4.28g (Elemental analysis: %(N)=4.409), The product obtained was a mixture of two polymers (E+E') (Figure 5).

ATR/FTIR: 3379 (ν_{N-H}), 2990 (ν_{Car-H}), 2928 (ν_{CH_2}), 2120 ($\nu_{C=C}$), 1646 ($\nu_{N=N}$) and 1610 ($\nu_{Car=Car}$) cm^{-1} .

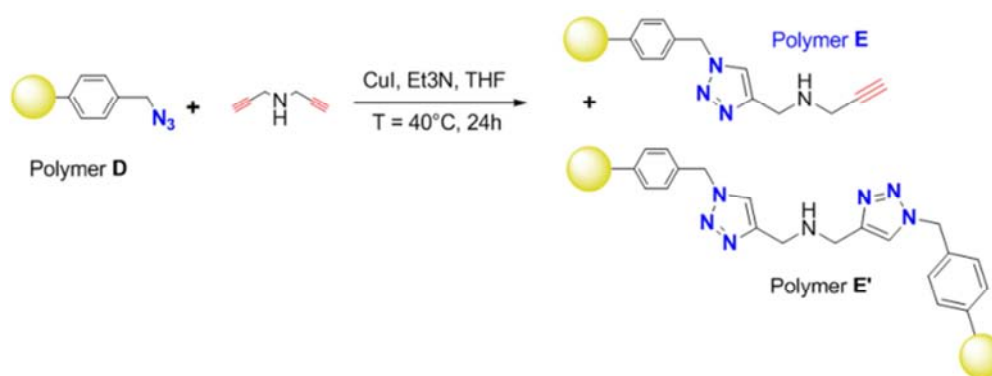


Figure 5. Reaction scheme of synthesis of [4-(N-propynyl amino methyl)triazol-1-yl]methyl polystyrene (E).

2.2.5. Synthesis of [4-{2-[2-(2-azidoethoxy)ethoxy]ethyl}triazol-1-yl-(N,N-dimethylamino)triazol-1-yl]methyl polystyrene (F)

The method of synthesis of polymer (F) was described in Figure 6. In 85mL of THF, 4g of the mixture (E+E') was slowly agitated for 30min at room temperature. 1-azido-2-[2-(2-azidoethoxy)ethoxy]ethane (A: 5.4 g, 27mmol), CuI (2mg, 0.01mmol) and Et₃N (5mL, 3.63g 36mmol) have been added. The suspension was slowly stirred

at 40°C for 72 hours. After filtration under vacuum, the obtained product was washed three times with THF (10mL), ET₃N (10mL), CH₂Cl₂ (10mL) and MeOH (10mL) alternatively and finally dried under vacuum at 40°C for 48 hours. The obtained product was a mixture of three polymers (F+E'+F'), the weight of the mixture of polymers was 3.54g (Elemental analysis: %(N) = 4.262).

ATR/FTIR: 3026 (ν_{Car-H}), 2911 (ν_{CH_2}), 2104 (ν_{N_3}), 1602 ($\nu_{N=N}$) cm^{-1} .

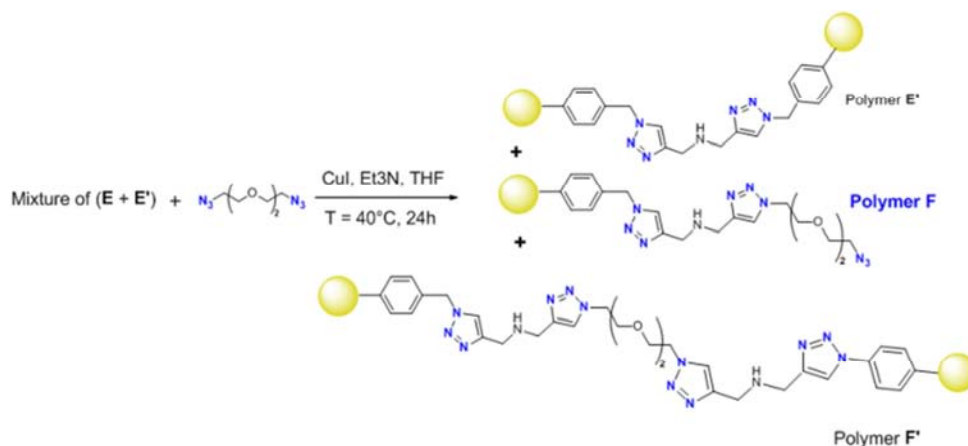


Figure 6. Reaction scheme of synthesis of [4-{2-[2-(2-azidoethoxy)ethoxy]ethyl}triazol-1-yl-(*N,N*-dimethylamino)triazol-1-yl]methyl polystyrene (F).

2.2.6. Synthesis of

[4-(2-{2-2-(4-[(1*S*,3*R*,5*R*,7*R*)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl]triazol-1-yl)ethoxy)ethoxy}ethyl)triazol-1-yl-(*N,N*-dimethylamino)triazol-1-yl]methylpolystyrene (G)

At 1g of the polymers (F+E'+F') in 10mL of DMF, we add 1g of (1*S*, 3*R*, 5*R*, 7*R*)-5,7-dihydroxy-2,2-diméthyl-N-(prop-2-yn-1-yl)hexahydro-1,3-benzodioxole-5-carboxamide (3.71mmol) (C), 2mg of cooper iodide (0.01mmol) and 1.5mL of ET₃N (1.22g,

12mmol), the mixture was stirred at 40°C for 24h. The obtained new product was filtered under vacuum and washed with DMF (10mL), ET₃N (10mL), CH₂Cl₂ (10mL), MeOH (10mL) three times alternatively and finally washed with distilled water (150mL). After drying at 60°C for two days under vacuum, the weighted of obtained product was 1g. The obtained product was a mixture of three polymers (G+E'+F'), (Figure 7).

ATR/FTIR: 3450 ($\nu_{\text{O-H}}$), 3034 ($\nu_{\text{Car-H}}$), 2937 (ν_{CH_2}), 1725 ($\nu_{\text{C=O}}$), 1602 ($\nu_{\text{N=N}}$) cm^{-1} .

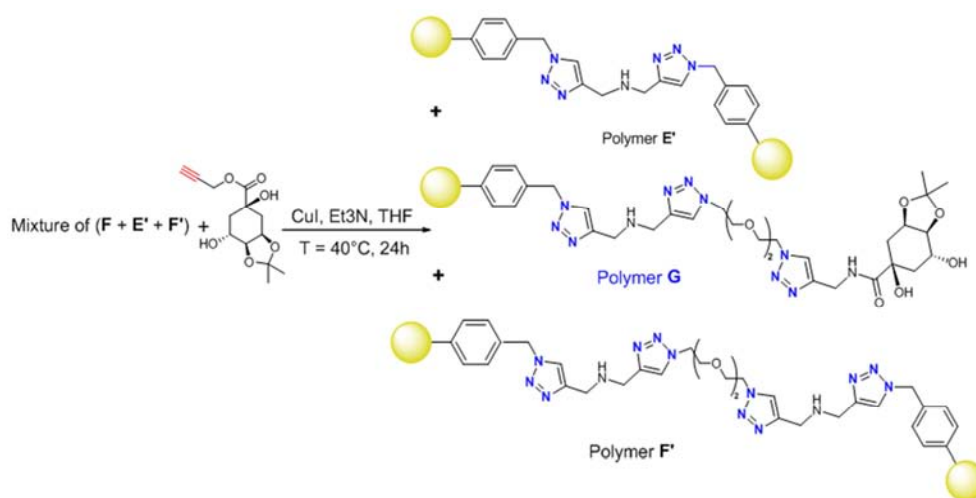


Figure 7. Reaction scheme of synthesis of [4-(2-{2-2-(4-[(1*S*,3*R*,5*R*,7*R*)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl]triazol-1-yl)ethoxy)ethoxy}ethyl)triazol-1-yl-(*N,N*-dimethylamino)triazol-1-yl]methylpolystyrene (G).

2.2.7. Synthesis of

[4-{(1*S*,3*R*,5*R*,7*R*)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl}triazol-1-yl]methylpolystyrene (H)

The polymer (H) presented in Figure 8 was synthesized by the condensation reaction between 1g (0.89mmol N₃) of polystyrene azide (D) and (1*S*, 3*R*, 5*R*, 7*R*)-5,7-dihydroxy-2,2-diméthyl-N-(prop-2-yn-1-yl)hexahydro-1,3-benzodioxole-5-carboxamide (C) (0.78g, 2.78mmol, 3.12eq) in presence of copper iodide (6mg,

0.03mmol, 3.4%) and 1.5mL of ET₃N (1.12g, 11mmol, 12.5eq) in 10mL of DMF. The suspension was stirred at 40°C during 72 hours, after filtration under vacuum, the polymer (H) was washed three times with DMF (10mL), ET₃N (10mL), CH₂Cl₂ (10mL), MeOH (10mL) alternatively and finally washed with 150mL of distilled water, this polymer was dried under vacuum at 60°C for two days. The weight of new isolated polymer was 0.9g, which give the yield 73%.

ATR/FTIR: 3433 ($\nu_{\text{O-H}}$), 3026 ($\nu_{\text{Car-H}}$), 2919 (ν_{CH_2}), 1735 ($\nu_{\text{C=O}}$), 1655 ($\nu_{\text{N=N}}$) cm^{-1} .

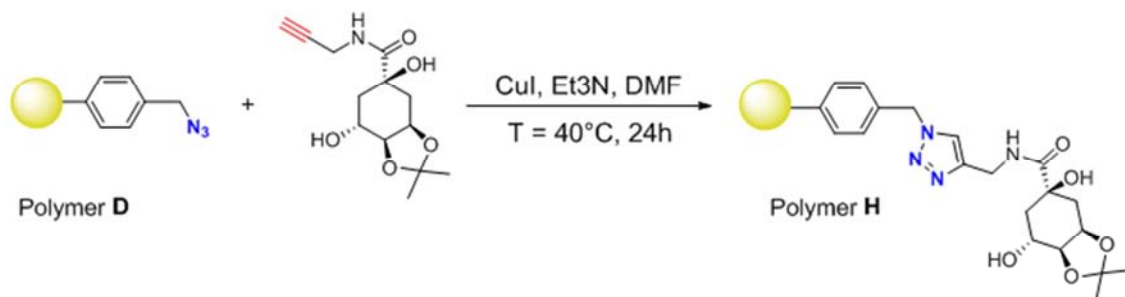


Figure 8. Reaction of synthesis of [4-{(1S,3R,5R,7R)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl}triazol-1-yl]methylpolystyrene (H).

2.2.8. Extraction of Zn and Mg Metals Located in Wastewater

The polymers (G+E'+F' & H) have been studied independently in the elimination of Zn and Mg located in wastewater. 0.2g of polymer was incubated with 20mL of wastewater at room temperature for 48 hours, the suspension was filtrated and washed with distilled water (5 x 50mL), after drying polymers, the solid phase was washed with 50mL of distilled water at T = 50°C during 20mn for desorption of metals (Zn & Mg) retained by studied polymers. After filtration, the obtained aqueous phase was analyzed by atomic absorption spectroscopy for dosing the amount of Zn and Mg liberated by polymers in distilled water.

3. Results and Discussion

Polymers synthesized in this work (Figures 4-8) were evaluated by IR spectroscopy (Figures 9 and 10), Differential Thermal Analysis DTA-TGA (Figure 11 and Table 1) and UV-Visible (Figure 12). Extraction results of Zn and Mg were presented in (Figures 13 and 14). Proposed mechanism of Zn and Mg extraction by studied polymers was presented in Figure 15.

3.1. Azidomethyl Polystyrene (D)

The percentage of nitrogen present in the polymer (D) was 3.726%, which gives 2,66mmol N.g⁻¹, for a final substitution of 0.89mmol N₃.g⁻¹, the yield of the substitution reaction estimated by elemental analysis was in the order of 81%. This substitution reaction was confirmed by IR spectroscopy (Figure 10), the intense band observed at 2095 cm⁻¹ was attributed for the vibrations of group N₃.

3.2. [4-(N-propynyl Amino Methyl)triazol-1-yl]methyl Polystyrene (E)

The coupling reaction between polymer (D) and di-propargylamine give a mixture of two polymers (E+E'). The proposed structure for (E) was confirmed by IR spectroscopy (Figure 9), but the structure of (E') was similar to that obtained by Sung *et al.* [17]. The percentage of nitrogen present in (E+E') was 4.409% or 2.9mmol of N.g⁻¹. The separation polymers (E+E') was practically difficult because they are not soluble in organic solvents. The IR spectrum of

(E+E') presented in (Figure 9) shows a low intensity band at 2120cm⁻¹ characterizing the vibrations of the alkyne function (C≡CH), disappearance of the band characterizing N₃ group observed at 2095cm⁻¹ in the Figure 10 (Polymer D), and appearance of new bands at 3344 and 1648cm⁻¹ characterizing respectively the vibrations of (NH) and (N=N) functions.

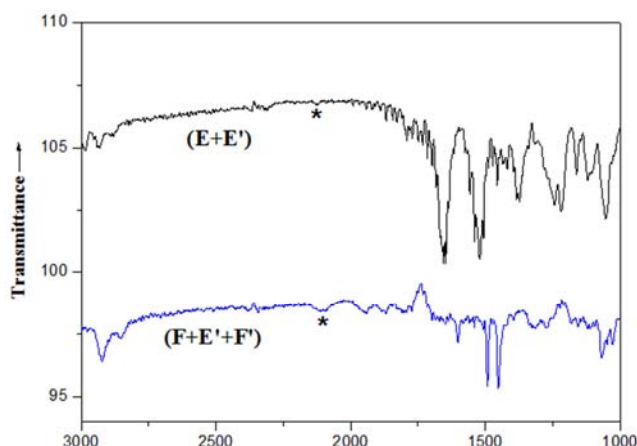


Figure 9. IR spectrum of polymers (E+E') and (F+E'+F').

3.3. [4-{2-[2-(2-azidoethoxy)ethoxy]ethyl}triazol-1-yl-(N,N-dimethylamino)Triazol-1-yl]Methyl Polystyrene (F)

The reaction between the mixture of polymers (E+E') and 1-azido-2-[2-(2-azidoethoxy)ethoxy]ethane (A) gives a mixture of polymers (F+E'+F') (Figure 6). The proposed structure of the polymer F was confirmed by IR spectroscopy presented in Figure 9, the proposed structure of F' was similar to that obtained by Chernykh *et al.* [18]. The percentage of nitrogen found in this mixture was 4.262% or 3.04mmol of N.g⁻¹. The separation of the polymers (F, E' & F') was practically difficult because they are not soluble in organic solvents.

If we look at the Figure 9, we observe an average band at 2104cm⁻¹ attributed to the vibrations of (N₃) group and a strong band observed at 1605cm⁻¹ attributed to the vibrations of (N=N) of the triazole heterocyclic, The amount of azide not reacted was low, but the amount of triazole formed was important, which explains the formation of the mixture consisting by the polymers (F+F'). Differential Thermal Analysis (DTA) presented in the Figure 11 shows one

endothermic peak at 412°C attributed for the decomposition of the grafted molecule on the Merrifield polystyrene, this decomposition was confirmed by a weight loss ($\Delta m = -7.88\text{mg}$, $\% = 78.80$) observed in the thermo gravimetric curve (Figure 11).

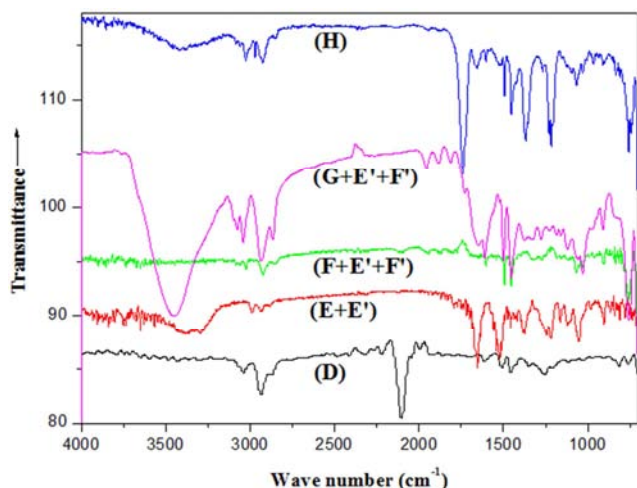


Figure 10. IR spectrum of synthesized polymers D to H.

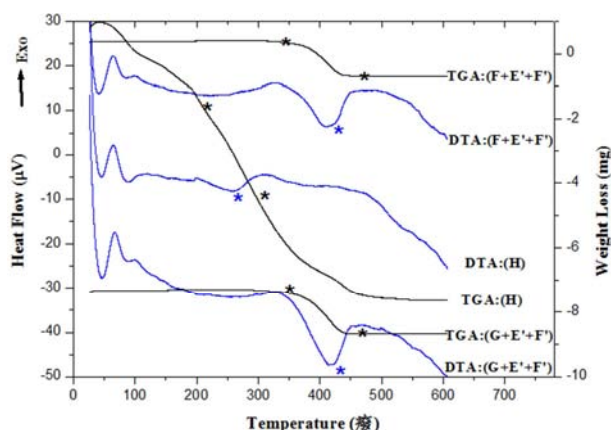


Figure 11. DTA-TGA curves for polymers D to H.

3.4. [4-(2-{2-(4-[(1S,3R,5R,7R)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl]triazol-1-yl)ethoxy}ethoxy)ethyl)triazol-1-yl-(N,N)-dimethylamino)triazol-1-yl]methylpolystyrene (G)

The coupled reaction between the alkyne (1S, 3R, 5R, 7R)-5,7-dihydroxy-2,2-diméthyl-N-(prop-2-yn-1-yl) hexahydro-1,3-benzodioxole-5-carboxamide (E) and polymers (F+E'+F'), give the mixture of polymers (G+E'+F'). The polymer (G) was confirmed by IR spectroscopy. IR spectrum presented in Figure 10 shows the disappearance of the band (N_3) observed at 2104cm^{-1} (Figure 9), the appearance of new bands at 3400 , 1750 and 1605cm^{-1} attributed to the vibration of (OH), (C=O) and (N=N), respectively.

Differential Thermal Analysis (DTA) and thermo gravimetric (TGA) presented in (Figure 11) shows an endothermic peak observed at 414°C attributed to its decomposition, this decomposition was accompanied by a weight loss ($\Delta m = -9.67\text{mg}$, $\% = 77.34$) (Figure 11).

3.5. [4-{(1S,3R,5R,7R)-5,7-dihydroxy-2,2-dimethylhexahydro-1,3-benzodioxole-5-carboxamidomethyl}triazol-1-yl]methylpolystyrene (H)

The coupled reaction between the alkyne (1S, 3R, 5R, 7R)-5,7-dihydroxy-2,2-diméthyl-N-(prop-2-yn-1-yl)hexahydro-1,3-benzodioxole-5-carboxamide (E) and polymer (D) catalyzed by cooper iodide give the polymer (H), structure of this polymer was confirmed by IR spectroscopy and thermal analysis (DTA-TGA). If we look at the IR spectrum (Figure 10), we observe the existence of characteristic bands at 3446 , 1732 and 1651cm^{-1} characterizing the vibrations of (OH), (C=O) and (N=N) groups, respectively. On the other hand, we observed the disappearance band of the N_3 function.

The DTA-TGA curves of the polymer (H) presented in Figure 11 show an endothermic peak at 256°C accompanied by a lost mass $\Delta m = -3.73\text{mg}$ ($\% = -25.73$). This lost mass corresponds to the decomposition of the organic molecule grafted on the polystyrene Merrifield, this results are consistent with these obtained from elemental analysis ($0.89\text{mmol N}_3\cdot\text{g}^{-1}$).

The Figure 12 present the UV-Vis spectra of the synthesized polymers realized in DMF, these results show the existence of an intense peak at 270nm attributed for the transition ($\pi \rightarrow \pi^*$), the polymer (H) present a second transition ($n \rightarrow \pi^*$) observed at 310nm , this transition was absent in the other polymers because their solubility in DMF was low. Results of UV-Vis spectroscopy are consistent with those found by Samanta *et al.* [19], which proves that synthesized polymers may fall in the category of semiconductors.

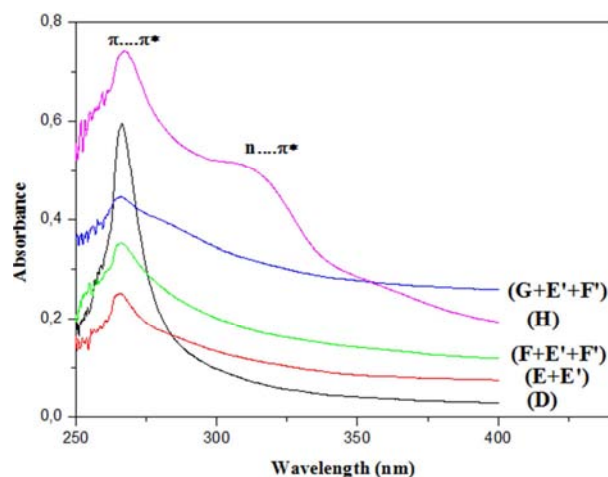


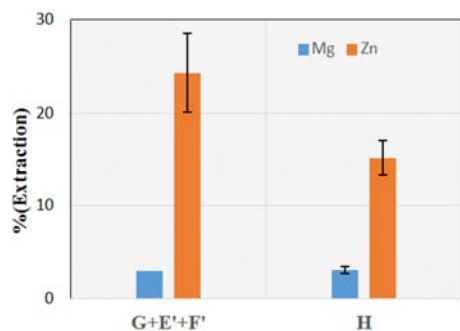
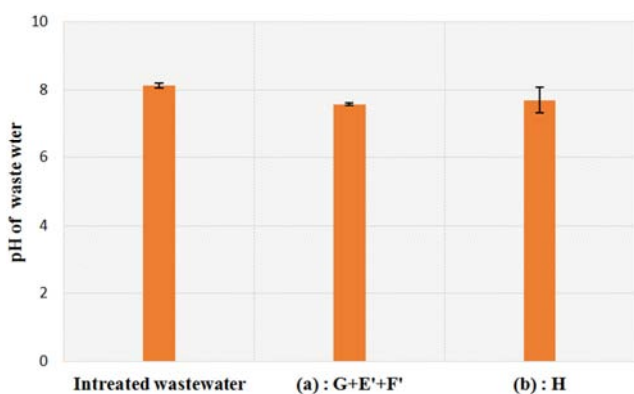
Figure 12. UV absorption spectra of polymers D to H realized in DMF solvent.

3.6. Extraction of Zn and Mg by Synthesized Polymers

We present in Figure 13 the percentage of extraction of Zn and Mg by polymers (G+E'+F') and (H), pH results of wastewater treated by (G+E'+F') and (H) are shown in Figure 14.

Table 1. DTA and TGA studies of synthesized polymers from polystyrene.

Polymer	Initial mass (mg)	T _{endo} (°C)	ΔH (J.g ⁻¹)	Δm (mg)	% (Δm)
F+E'+F'	10	412	575.20	-7.88	-78.80
G+E'+F'	12.5	414	579.76	-9.67	-77.36
H	14	256	143.07	-3.732	-26.65

**Figure 13.** % (Extraction) of Zn and Mg obtained after treatment wastewater by (G+E'+F') and (H).**Figure 14.** Ph of wastewater before and after treatment (a) by (G+E'+F') and (b) by (H).

The concentrations of metals located in waste water of Beja station (Tunisia) before treatment by studied polymers were presented in Table 2.

Table 2. Concentration of metals localized in untreated wastewater of Beja station.

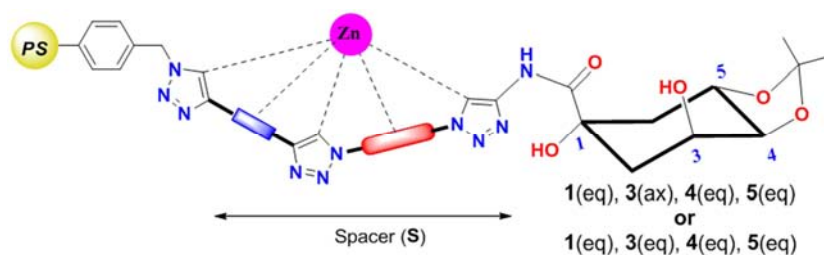
Metal	Mg	Cu	Fe	Mn	Zn
Concentration (mg.L ⁻¹)	17.45	0.01	0.061	5.10 ⁻³	0.175

If we look at the Figure 13, we find that the polymers (G+E'+F') and (H) gave practically the similar percentages of extraction for Mg, the average percentage was 2.92 ± 0.20 . In this case, neither spacer (S) presented in Figure 1 nor grafted quinic acid derivative have any effect in the extraction of Mg. The polymers studied have no remarkable selectivity for the extraction of Mg.

On the other hand, the best removal percentage Zinc was obtained by the polymer (G+E'+F'), the value found was in order of $24.31 \pm 4.17\%$. The extraction percentage of Zn by (G+E'+F') was better than those found in the literature [20, 21]. Polymer (H) gave a lower percentage than that obtained using (G+E'+F'), it's in the order of $15.20 \pm 1.86\%$.

If we compare the structures of polymers (G+E'+F') and (H), we find that the terminal molecule of quinic acid was common for the two polymers, but the spacer (S) shown in Figure 1 was present only in (G+E'+F') and not in (H). The spacer (S) is responsible for the extraction of Zn, but the derivative of quinic acid has no effect on the extraction of metals because the oxygen atoms of quinic acid derivative are distant (Figure 15).

The pH of intreated and treated wastewater by polymers (G+E'+F') and (H) synthesized from polystyrene are presented in Figure 14. The pH of the intreated wastewater was about of 8.15. After treatment by polymers (G+E'+F') and (H), the pH decreases, the pH values for treated wastewater by (G+E'+F') and (H) obtained are 7.58 and 7.71 respectively. This decrease in pH can be explained by the extraction of several basic chemical species: nitrates, sulphates, phosphates and amines [22-24] or by exchange of H⁺ ions of quinic acid derivative with other cations [25].

**Figure 15.** Proposed mechanism for Zn extraction by synthesized polymers.

4. Conclusion

In this work, we have synthesized a new compounds from Merrifield polystyrene. These obtained compounds are a mixture of polymers containing (1,2,3)-triazole units. The separation of these polymers was practically impossible because they are not soluble in organic solvents. These polymers have been characterized by IR spectroscopy and

thermal analysis (DTA & TGA). These polymers have been tested in the extraction of Zn and Mg located in wastewater from Beja region (Tunisia). The obtained polymers do not have a selectivity for Mg, the percentage of extraction obtained did not exceed 3%. For Zn, the percentage obtained is 25% with the polymer (G+E'+F'). The percentage of extraction metals was low because other elements (nitrates, sulphates, phosphates and amines) have been removed by these polymers.

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